

Valences of Some Common Monatomic Ions

+1

Copper (I) Cu^+

Cesium Cs^+

Hydrogen H^+

Lithium Li^+

Potassium K^+

Silver Ag^+

Sodium Na^+

+2

Barium Ba^{+2}

Beryllium Be^{+2}

Cadmium Cd^{+2}

Calcium Ca^{+2}

Chromium (II) Cr^{+2}

Cobalt (II) Co^{+2}

Copper (II) Cu^{+2}

Iron (II) Fe^{+2}

Lead (II) Pb^{+2}

Magnesium Mg^{+2}

Manganese Mn^{+2}

Mercury (II) Hg^{+2}

Nickel Ni^{+2}

Strontium Sr^{+2}

Tin (II) Sn^{+2}

Zinc Zn^{+2}

+3

Aluminum Al^{+3}

Chromium (III) Cr^{+3}

Iron (III) Fe^{+3}

+4

Lead (IV) Pb^{+4}

Tin (IV) Sn^{+4}

+6

Chromium (VI) Cr^{+6}

-1

Bromide Br^{-1}

Chloride Cl^{-1}

Fluoride F^{-1}

Hydride H^{-1}

Iodide I^{-1}

-2

Oxide O^{-2}

Sulfide S^{-2}

Valences of Some Common Polyatomic Ions

+1

Ammonium NH_4^{+1}

-1

Acetate $\text{C}_2\text{H}_3\text{O}_2^{-1}$

Bromate BrO_3^{-1}

Chlorate ClO_3^{-1}

Chlorite ClO_2^{-1}

Cyanide CN^{-1}

Dihydrogen Phosphate $\text{H}_2\text{PO}_4^{-2}$

Hydrogen carbonate or

Bicarbonate HCO_3^{-1}

Hydrogen sulfate HSO_4^{-1}

Hydrogen sulfide HS^{-1}

Hydroxide OH^{-1}

Hypochlorite ClO^{-1}

Iodate IO_3^{-1}

Nitrate NO_3^{-1}

Nitrite NO_2^{-1}

Perchlorate ClO_4^{-1}

Permanganate MnO_4^{-1}

Thiocyanate SCN^{-1}

+2

Mercury (I) Hg_2^{+2}

-2

Carbonate CO_3^{-2}

Chromate CrO_4^{-2}

Dichromate $\text{Cr}_2\text{O}_7^{-2}$

Hydrogen Phosphate HPO_4^{-1}

Oxalate $\text{C}_2\text{O}_4^{-2}$

Peroxide O_2^{-2}

Silicate SiO_4^{-2}

Sulfate SO_4^{-2}

Sulfite SO_3^{-2}

Tartrate $\text{C}_4\text{H}_4\text{O}_6^{-2}$

Tetraborate $\text{B}_4\text{O}_7^{-2}$

Thiosulfate $\text{S}_2\text{O}_3^{-2}$

-3

Phosphate PO_4^{-3}