

Valences of Some Common Monatomic Ions

Common Monatomic Cations

Valence	Name	Ion
1+	Cesium	Cs ⁺
	Hydrogen	H ⁺
	Lithium	Li ⁺
	Potassium	K ⁺
	Sodium	Na ⁺
	Silver	Ag ⁺
2+	Barium	Ba ²⁺
	Beryllium	Be ²⁺
	Cadmium	Cd ²⁺
	Calcium	Ca ²⁺
	Magnesium	Mg ²⁺
	Manganese	Mn ²⁺
	Nickel	Ni ²⁺
	Strontium	Sr ²⁺
	Zinc	Zn ²⁺
3+	Aluminum	Al ³⁺

Common Monatomic Anions

Valence	Name	Ion
1-	Bromide	Br ⁻
	Chloride	Cl ⁻
	Fluoride	F ⁻
	Hydride	H ⁻
	Iodide	I ⁻
2-	Oxide	O ²⁻
	Sulfide	S ²⁻
3-	Nitride	N ³⁻
	Phosphide	P ³⁻

Common Type II Cations

Name	Ion
Chromium(II)	Cr ²⁺
Chromium(III)	Cr ³⁺
Chromium(VI)	Cr ⁶⁺
Cobalt(II)	Co ²⁺
Cobalt(III)	Co ³⁺
Copper(I)	Cu ⁺
Copper(II)	Cu ²⁺
Iron(II)	Fe ²⁺
Iron(III)	Fe ³⁺
Lead(II)	Pb ²⁺
Lead(IV)	Pb ⁴⁺
Mercury(I)	Hg ₂ ²⁺
Mercury(II)	Hg ²⁺
Tin(II)	Sn ²⁺
Tin(IV)	Sn ⁴⁺

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Common Polyatomic Ions

Name	Ion
Ammonium	NH_4^+
Mercury(I)	Hg_2^{2+}
Acetate	$\text{C}_2\text{H}_3\text{O}_2^-$
Bromate	BrO_3^-
Carbonate	CO_3^{2-}
Hydrogen Carbonate (or Bicarbonate)	HCO_3^-
Hypochlorite	ClO^-
Chlorite	ClO_2^-
Chlorate	ClO_3^-
Perchlorate	ClO_4^-
Dichromate	$\text{Cr}_2\text{O}_7^{2-}$
Chromate	CrO_4^{2-}
Thiocyanate	SCN^-
Cyanide	CN^-
Hydroxide	OH^-

Iodate	IO_3^-
Nitrite	NO_2^-
Nitrate	NO_3^-
Oxalate	$\text{C}_2\text{O}_4^{2-}$
Permanganate	MnO_4^-
Peroxide	O_2^{2-}
Phosphate	PO_4^{3-}
Hydrogen Phosphate	HPO_4^{2-}
Dihydrogen Phosphate	H_2PO_4^-
Silicate	SiO_4^{2-}
Sulfite	SO_3^{2-}
Sulfate	SO_4^{2-}
Hydrogen Sulfate (or bisulfate)	HSO_4^-
Hydrogen Sulfide	HS^-
Thiosulfate	$\text{S}_2\text{O}_3^{2-}$
Tartate	$\text{C}_4\text{H}_4\text{O}_6^{2-}$
Tetraborate	$\text{B}_4\text{O}_7^{2-}$